

KRISTU JYOTI VIDYA NIKETAN

CHETHIPUZHA, CHANGANASSERY – 686 104

Second Mid-Term Examination – 2022 – 23

Class : VIII

CHEMISTRY

Time : 1h

Max. Marks : 40

Section - A

(Compulsory)

Question 1

a) Choose the correct answers

5

- i. The chemical formula of calcium nitride is ---
A- CaNO_3 B- Ca_3N_2
C- CaNO_2 D- $\text{Ca}(\text{NO}_3)_2$

- ii. Which of the following does **not** have valency 2?
A- Calcium B- Potassium
C- Nickel D- Barium

- iii. Which of the following exhibit variable Valency?
A- Oxygen B- Zinc
C- Tin D- sodium

- iv. Which of the following will be able to displace magnesium Metal from magnesium sulphate?
A- copper B- calcium
C- iron D- zinc

- v. During which of the following reactions salt and water is formed as the only product?
A- Exothermic reaction B- decomposition reaction
C- neutralization reaction D- displacement reaction

b) Name the following

4

- i. A substance which increases the efficiency of a catalyst.
- ii. A reaction between two compounds in aqueous solution state to give a new compound one of which is insoluble.
- iii. The shorthand representation of an element
- iv. The Latin name for the element mercury.

- c) Give reason for the following 2
- i. A chemical equation should be balanced.
 - ii. No reaction is observed when copper metal is put in iron sulphate solution .

- d) Match the following 2

Column 1	Column 2
i. Element with valency zero	a. Silver
ii. An Anion	b. Neon
iii. An acidic oxide	c. Sulphate
iv. An element with variable valency	d. Sulphur dioxide

- e) Write the Chemical formula of the of the following atoms 4

- i. Calcium bicarbonate
- ii. Sodium bromide
- iii. Potassium nitrate
- iv. Aluminium hydroxide

- f) Name the following compounds 3

- i. $\text{Fe}_2(\text{SO}_4)_3$
- ii. ZnO
- iii. Na_2SO_3

Section B

Answer all questions

Question 2

- a) Classify the following reaction into Precipitation , Decomposition , Synthesis , Displacement , Neutralization. 5

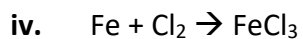
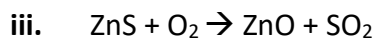
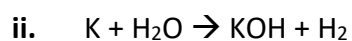
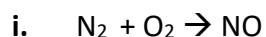
- i. $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- ii. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- iii. $2 \text{Zn} + \text{O}_2 \rightarrow 2 \text{ZnO}$
- iv. $\text{PbCO}_3 \rightarrow \text{PbO} + \text{CO}_2$
- v. $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl} \downarrow$

- b) Answer the following questions

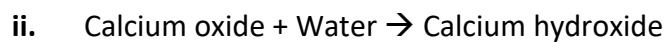
- i. Define double decomposition reaction 1
- ii. What is a catalyst? give an example. $1\frac{1}{2}$
- iii. What is an activity series? Name the element at the top of the activity series. $1\frac{1}{2}$
- iv. Differentiate between a photochemical decomposition and electrolytic decomposition. 1

Question 3

a) Balance the following equation. 4



b) Write balanced molecular equation for the word equation 2



c) Match the following 4

List A	List B
1. Iron(III)Sulphate	A. Cu_2O
2. Phosphoric acid	B. Cu_2S
3. Copper(II)Sulphide	C. H_3PO_4
4. Copper(I)Oxide	D. $FeSO_4$
	E. CuS
	F. $Fe_2(SO_4)_3$

Prepared by : Reena Mary Sunil